I. Fill in the blanks with the most appropriate term:

A ______ chemical ______ equation ______ tells the story of a chemical reaction. ______ Reactants ______ are the starting substances in the reaction while ______ products ______ are the new substances that are formed. The large numbers in front of some of the formulas are called ______ coefficients ______. These numbers are used to ______ balance ______ the equation because chemical reactions must obey the Law of ______ Conservation ______ of Matter. The number of atoms of each element on both sides of the equation must be ______ equal ______ because matter cannot be ______ created ______ or ______ destroyed ______. When balancing equations, the only numbers that can be changed are ______ coefficients ______; remember that ______ formulas / subscripts ______ must never be changed in order to balance an equation.

II. Balance the following equations:

1. \( 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3 \)

2. \( \text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O} \)

3. \( \text{Al(NO}_3)_3 + 3\text{NaOH} \rightarrow \text{Al(OH)}_3 + 3\text{NaNO}_3 \)
4. $2\text{KNO}_3 \rightarrow 2\text{KNO}_2 + \text{O}_2$

5. $3\text{O}_2 + \text{CS}_2 \rightarrow \text{CO}_2 + 2\text{SO}_2$

6. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

7. $3\text{BaF}_2 + 2\text{K}_3\text{PO}_4 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6\text{KF}$

8. $\text{H}_2\text{SO}_4 + \text{Mg(NO}_3)_2 \rightarrow \text{MgSO}_4 + 2\text{HNO}_3$

9. $2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2$

10. $\text{WO}_3 + 3\text{H}_2 \rightarrow \text{W} + 3\text{H}_2\text{O}$